

CANDIDATE  
NAME

CENTRE  
NUMBER

--	--	--	--	--

CANDIDATE  
NUMBER

--	--	--	--



**CHEMISTRY**

**5070/31**

Paper 3 Practical Test

**October/November 2017**

**1 hour 30 minutes**

Candidates answer on the Question Paper.

Additional Materials: As listed in the Confidential Instructions

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

Electronic calculators may be used.

Qualitative Analysis Notes are printed on page 8.

You should show the essential steps in any calculations and record experimental results in the spaces provided on the Question Paper.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
<b>Total</b>	

This document consists of **6** printed pages and **2** blank pages.

- 1 The composition of concentrated nitric acid can be determined by diluting a sample of the acid and then titrating aqueous sodium carbonate of known concentration with the diluted acid.

**P** is dilute nitric acid. It has been prepared by adding  $10.0\text{ cm}^3$  of the concentrated nitric acid to distilled water and then making the total volume of the solution up to  $500\text{ cm}^3$  in a volumetric flask by adding distilled water.

**Q** is  $0.153\text{ mol/dm}^3$  sodium carbonate.

- (a) Put **P** into the burette.

Pipette a  $25.0\text{ cm}^3$  (or  $20.0\text{ cm}^3$ ) portion of **Q** into a flask and titrate with **P**, using the indicator provided.

Record your results in the table, repeating the titration as many times as you consider necessary to achieve consistent results.

### Results

#### *Burette readings*

titration number	1	2	
final reading / $\text{cm}^3$			
initial reading / $\text{cm}^3$			
volume of <b>P</b> used / $\text{cm}^3$			
best titration results (✓)			

### Summary

Tick (✓) the best titration results.

Using the best titration results, the average volume of **P** required was .....  $\text{cm}^3$ .

Volume of **Q** used was .....  $\text{cm}^3$ .

[12]

- (b) **Q** is 0.153 mol/dm<sup>3</sup> sodium carbonate.

Using your results from (a), calculate the concentration, in mol/dm<sup>3</sup>, of nitric acid in **P**. Give your answer to three significant figures.



concentration of nitric acid in **P** ..... mol/dm<sup>3</sup> [2]

- (c) Using your answer from (b), calculate the number of moles of nitric acid in 10.0 cm<sup>3</sup> of concentrated nitric acid.

moles of nitric acid in 10.0 cm<sup>3</sup> of concentrated nitric acid ..... [1]

- (d) Using your answer from (c), calculate the concentration, in mol/dm<sup>3</sup>, of concentrated nitric acid.

concentration of concentrated nitric acid ..... mol/dm<sup>3</sup> [1]

- (e) Using your answer from (d), calculate the mass, in g, of nitric acid in 1 dm<sup>3</sup> of concentrated nitric acid.  
[The relative formula mass of nitric acid is 63.]

mass of nitric acid in 1 dm<sup>3</sup> of concentrated nitric acid ..... g [1]

[Total: 17]

2 You are provided with two solutions, **R** and **S**.

(a) Carry out the following tests and record your observations in the table.

You should test and name any gas evolved.

test no.	test	observations with solution <b>R</b>	observations with solution <b>S</b>
<b>1</b>	<p>(a) To 2 cm depth of the solution in a test-tube, add aqueous sodium hydroxide until a change is seen.</p> <p>(b) To the mixture from (a), add excess aqueous sodium hydroxide.</p>		
<b>2</b>	<p>(a) To 2 cm depth of the solution in a test-tube, add aqueous ammonia until a change is seen.</p> <p>(b) To the mixture from (a), add excess aqueous ammonia. Keep the final mixture for use in (c).</p> <p>(c) To 1 cm depth of aqueous hydrogen peroxide in a boiling tube, add the final mixture from (b).</p>		

test no.	test	observations with solution <b>R</b>	observations with solution <b>S</b>
<b>3</b>	<p><b>(a)</b> To 1 cm depth of the solution in a test-tube, add an equal volume of dilute nitric acid.</p> <p><b>(b)</b> Pour half of the mixture from <b>(a)</b> into a test-tube and add an equal volume of aqueous barium nitrate.</p> <p><b>(c)</b> To the other half of the mixture from <b>(a)</b>, add an equal volume of aqueous silver nitrate.</p>		

[21]

**(b) Conclusions**Identify the compound in solution **R**.The compound in solution **R** is .....Identify the compound in solution **S**.The compound in solution **S** is .....

[2]

[Total: 23]





## QUALITATIVE ANALYSIS NOTES

### Tests for anions

<i>anion</i>	<i>test</i>	<i>test result</i>
carbonate ( $\text{CO}_3^{2-}$ )	add dilute acid	effervescence, carbon dioxide produced
chloride ( $\text{Cl}^-$ ) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
iodide ( $\text{I}^-$ ) [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	yellow ppt.
nitrate ( $\text{NO}_3^-$ ) [in solution]	add aqueous sodium hydroxide, then add aluminium foil; warm carefully	ammonia produced
sulfate ( $\text{SO}_4^{2-}$ ) [in solution]	acidify with dilute nitric acid, then add aqueous barium nitrate	white ppt., insoluble in excess dilute nitric acid

### Tests for aqueous cations

<i>cation</i>	<i>effect of aqueous sodium hydroxide</i>	<i>effect of aqueous ammonia</i>
aluminium ( $\text{Al}^{3+}$ )	white ppt., soluble in excess giving a colourless solution	white ppt., insoluble in excess
ammonium ( $\text{NH}_4^+$ )	ammonia produced on warming	–
calcium ( $\text{Ca}^{2+}$ )	white ppt., insoluble in excess	no ppt.
chromium(III) ( $\text{Cr}^{3+}$ )	green ppt., soluble in excess, giving a green solution	green ppt., insoluble in excess
copper(II) ( $\text{Cu}^{2+}$ )	light blue ppt., insoluble in excess	light blue ppt., soluble in excess giving a dark blue solution
iron(II) ( $\text{Fe}^{2+}$ )	green ppt., insoluble in excess	green ppt., insoluble in excess
iron(III) ( $\text{Fe}^{3+}$ )	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc ( $\text{Zn}^{2+}$ )	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

### Tests for gases

<i>gas</i>	<i>test and test result</i>
ammonia ( $\text{NH}_3$ )	turns damp red litmus paper blue
carbon dioxide ( $\text{CO}_2$ )	turns limewater milky
chlorine ( $\text{Cl}_2$ )	bleaches damp litmus paper
hydrogen ( $\text{H}_2$ )	'pops' with a lighted splint
oxygen ( $\text{O}_2$ )	relights a glowing splint

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge International Examinations Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cie.org.uk](http://www.cie.org.uk) after the live examination series.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.