

Cambridge International Examinations

Cambridge International General Certificate of Secondary Education

0620/11 **CHEMISTRY**

October/November 2016 Paper 1 Multiple Choice (Core)

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO **NOT** WRITE IN ANY BARCODES.

There are forty questions on this paper. Answer all questions. For each question there are four possible answers A, B, C and D.

Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of 17 printed pages and 3 blank pages.



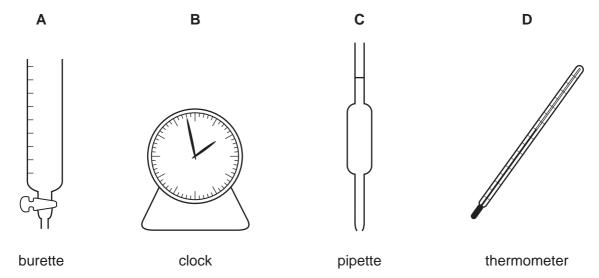
1 'Particles moving **very slowly** from an area of higher concentration to an area of lower concentration.'

Which process is being described?

- A a liquid being frozen
- B a solid melting
- **C** a substance diffusing through a liquid
- **D** a substance diffusing through the air
- 2 A student mixes 25 cm³ samples of dilute hydrochloric acid with different volumes of aqueous sodium hydroxide.

In each case, the student measures the change in temperature to test if the reaction is exothermic.

Which piece of apparatus is **not** needed?



3 Information about the solubility of four solids, P, Q, R and S, is given in the table.

	Р	Q	R	S
solubility in water	dissolves	insoluble	insoluble	dissolves

A student attempted to separate mixtures of these solids using the following method.

- 1 Add the mixture to a beaker of water and stir.
- 2 Filter the mixture.
- 3 Crystallise one of the solids from the filtrate.

Which of the following mixtures could **not** be separated by this method?

- A a mixture of P and R
- B a mixture of Q and P
- C a mixture of Q and R
- **D** a mixture of R and S

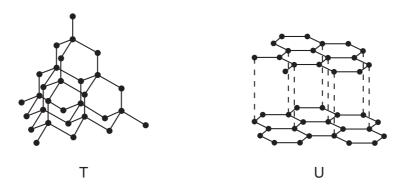
4 The table shows information about atoms of three different elements.

element	proton number	nucleon number	number of protons	number of neutrons	number of electrons
chlorine	17	35	17	W	17
chlorine	17	Х	17	19	17
argon	Y	40	18	22	18
potassium	19	39	19	20	Z

What are the values of W, X, Y and Z?

	W	X	Υ	Z
Α	18	35	18	19
В	18	36	18	19
С	19	35	19	18
D	19	36	19	18

5 The diagrams show the structures of two forms of the same element.



What are the reasons for using T in cutting tools and U as a lubricant?

	Т	U
Α	It is very hard because each atom is held in place by strong covalent bonds.	The layers can slide over each other because the covalent bonds are weak.
В	It is very hard because each atom is held in place by strong covalent bonds.	The layers can slide over each other due to weak forces between the layers.
С	It is very hard because there are no electrons able to move.	The layers can slide over each other because the covalent bonds are weak.
D	It is very hard because there are no electrons able to move.	The layers can slide over each other due to weak forces between the layers.

6 Ions are formed by elements losing or gaining electrons.

Which statement is correct?

- **A** Metal atoms gain electrons to form positive ions.
- **B** Non-metal atoms lose electrons to form positive ions.
- **C** The charge on an ion is always either +1 or -1.
- **D** Group I ions have the same electronic structure as noble gases.
- 7 A molecule of X contains two carbon atoms, four hydrogen atoms and two oxygen atoms.

What is the formula of X?

A CH₂CO₂H B CH₃COH C CH₃COOH D C₂H₃COOH

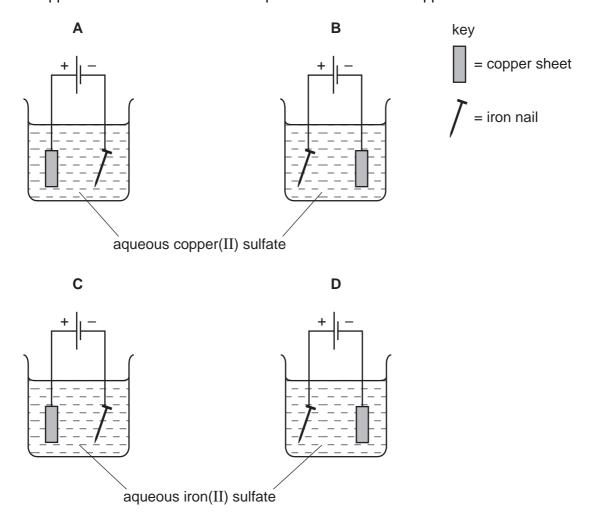
8 Concentrated aqueous potassium chloride is electrolysed using platinum electrodes.

The solution contains the ions K^+ , Cl^- , H^+ and OH^- .

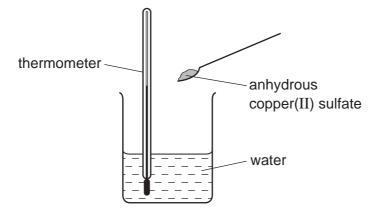
Which electrodes are the ions attracted to during this electrolysis?

	anode	cathode
Α	C <i>l</i> ⁻ and K⁺	H⁺ and OH⁻
В	C <i>l</i> ⁻ and OH⁻	H⁺ and K⁺
С	H⁺ and K⁺	C <i>l</i> ⁻ and OH⁻
D	H⁺ and OH⁻	C <i>l</i> ⁻ and K⁺

9 Which apparatus could be used to electroplate an iron nail with copper?



10 When anhydrous copper(II) sulfate is added to water a solution is formed and heat is given out.



Which row shows the temperature change and the type of reaction taking place?

	temperature change	type of reaction
Α	decrease	endothermic
В	decrease	exothermic
С	increase	endothermic
D	increase	exothermic

11 The combustion of element X releases large amounts of energy.

What is X?

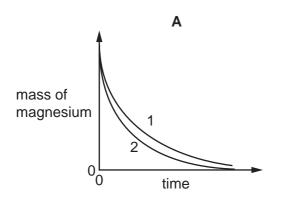
- A ethanol
- **B** hydrogen
- C methane
- **D** uranium

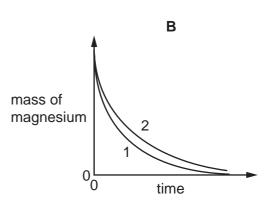
12 The rate of reaction between magnesium and excess dilute hydrochloric acid was followed by measuring the mass of magnesium present at regular time intervals.

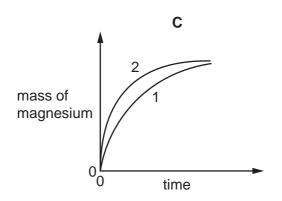
Two experiments were performed.

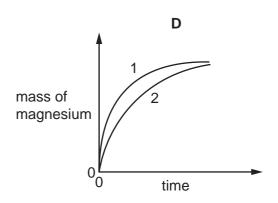
Both experiments used 0.1 g of magnesium ribbon. The acid in experiment 1 was less concentrated than in experiment 2.

Which graph shows the results of the experiments?









13 Which reaction is reversible?

A
$$CuCO_3 + 2HCl \rightarrow CuCl_2 + CO_2 + H_2O$$

B
$$CuSO_4.5H_2O \rightarrow CuSO_4 + 5H_2O$$

C
$$2Na + 2H_2O \rightarrow 2NaOH + H_2$$

D NaOH + HC
$$l \rightarrow$$
 NaC $l +$ H₂O

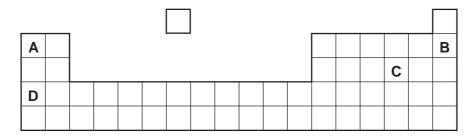
14 Tin is formed when tin(II) oxide is heated with carbon.

What happens to the tin in the tin(II) oxide in this reaction?

- A It is burnt.
- **B** It is electrolysed.
- C It is oxidised.
- **D** It is reduced.

15 Part of the Periodic Table is shown.

Which element forms an acidic oxide?



16 Four substances, P, Q, R and S, are tested as shown.

toot	substance				
test	Р	Q	R	S	
dilute hydrochloric acid added	gas given off which 'pops' with a lighted splint	gas given off which turns limewater milky	no reaction	no reaction	
dilute aqueous sodium hydroxide added and warmed gently	no reaction	no reaction	gas given off which turns damp, red litmus paper blue	no reaction	

What are P, Q, R and S?

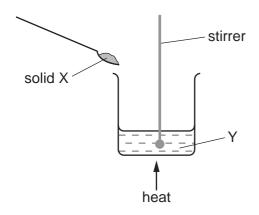
	Р	Q	R	S
Α	Mg	Na ₂ CO ₃	NH₄C <i>l</i>	NaC <i>l</i>
В	Mg	NH₄C <i>l</i>	Na ₂ CO ₃	NaC <i>l</i>
С	Mg	Na ₂ CO ₃	NaC <i>l</i>	NH₄C <i>l</i>
D	Na ₂ CO ₃	Mg	NaC <i>l</i>	NH₄C <i>l</i>

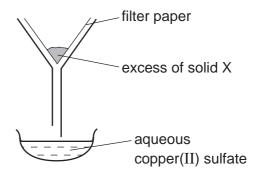
17 Acids can react with metal oxides, carbonates and metals.

Which reactions produce a gas?

	acid with metal oxide	acid with carbonate	acid with metal	
Α	✓	✓	✓	key
В	✓	X	X	√ = gas is produced
С	X	✓	✓	X = no gas is produced
D	×	✓	×	

18 The apparatus shown is used to prepare aqueous copper(II) sulfate.





What are X and Y?

	Х	Υ	
Α	copper	aqueous iron(II) sulfate	
В	copper(II) chloride	sulfuric acid	
С	copper(II) oxide	sulfuric acid	
D	sulfur	aqueous copper(II) chloride	

- 19 Which statement about trends in the Periodic Table is **not** correct?
 - **A** Elements in the same period have the same number of electron shells.
 - **B** The elements change from metals to non-metals from left to right.
 - **C** The number of protons in an atom of an element increases from left to right.
 - **D** The oxides of the elements change from acidic to basic from left to right.
- 20 What is **not** a property of Group I metals?
 - **A** They are soft and can be cut with a knife.
 - **B** They react when exposed to oxygen in the air.
 - **C** They produce an acidic solution when they react with water.
 - **D** They react rapidly with water producing hydrogen gas.
- 21 Which statement about the element with proton number 54 is correct?
 - A It burns in the air to form an oxide.
 - **B** It could be used in balloons because it has a very low density.
 - **C** It is a gas at room temperature.
 - **D** It is reactive because it has a full outer shell of electrons.

22 Which element is a transition element?

	colour of chloride	melting point of element/°C
Α	orange	113
В	orange	1535
С	white	113
D	white	1535

23 Which row describes the trends in the properties of the Group VII elements as the group is descended?

	colour	density	reactivity with halide ions
Α	darkens	decreases	increases
В	darkens	increases	decreases
С	lightens	decreases	increases
D	lightens	increases	decreases

24 Four metals are listed in decreasing order of reactivity.

magnesium

zinc

iron

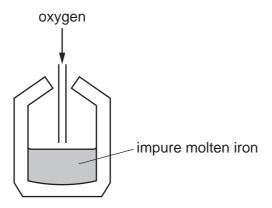
copper

Titanium reacts with acid and cannot be extracted from its ore by heating with carbon.

Where should titanium be placed in the list?

- A below copper
- **B** between iron and copper
- C between magnesium and zinc
- **D** between zinc and iron

25 Impure iron from the blast furnace is converted to steel as shown.



Which statement about the process is correct?

- A Acidic oxides are added to remove alkaline impurities.
- **B** Coke is added as a reducing agent.
- **C** Oxygen is blown in to oxidise the impure iron.
- **D** The steel produced contains less carbon than the impure iron.
- 26 A student added dilute hydrochloric acid to four metals and recorded the results.

Some of the results are not correct.

	res	ults	
	metal	gas given off yes yes no yes	
1	copper	yes	
2	iron	yes	
3	magnesium	no	
4	zinc	yes	

Which two results are correct?

- 1 and 3 **B** 1 and 4
 - **C** 2 and 3
- **D** 2 and 4

27 Some properties of three metals, P, Q and R, are shown.

metal	density	resistance to corrosion	electrical conductivity
Р	low	high	very good
Q	high	high	very good
R	low	low	good

Which metals would be suitable for use in electrical wiring and aircraft manufacture?

	electrical wiring	aircraft manufacture
Α	Р	Q
В	Q	Р
С	Q	R
D	R	Р

28 One sample of sea-water is distilled while another sample of sea-water is filtered.

Which statement about the samples is correct?

- **A** The distilled sample boils at exactly 100 °C and contains dissolved salts.
- **B** The distilled sample boils at 103 °C and does **not** contain dissolved salts.
- **C** The filtered sample boils at 103 °C and contains dissolved salts.
- **D** The filtered sample boils at exactly 100 °C and does **not** contain dissolved salts.
- **29** Air is a mixture of gases.

Which gas is present in the largest amount?

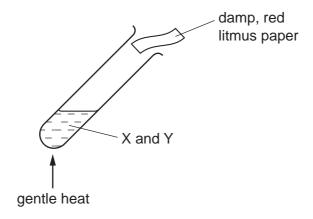
- **A** argon
- **B** carbon dioxide
- **C** nitrogen
- **D** oxygen

30 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane	
Α	formed when vegetation decomposes	✓	X	key
В	greenhouse gas	✓	✓	✓ = true
С	present in unpolluted air	×	×	x = false
D	produced during respiration	X	✓	

31 A mixture of two substances, X and Y, is heated.

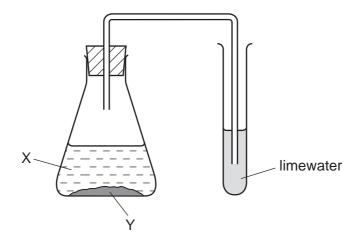
The damp, red litmus paper turns blue.



What are X and Y?

	X	Υ
Α	aluminium nitrate	hydrochloric acid
В	aluminium nitrate	sodium hydroxide solution
С	ammonium chloride	hydrochloric acid
D	ammonium chloride	sodium hydroxide solution

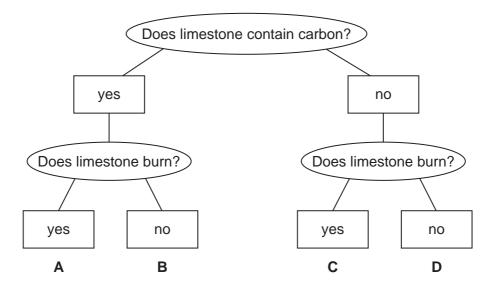
32 In the experiment shown, a white precipitate forms in the limewater.



What are X and Y?

	X	Υ
Α	aqueous sodium hydroxide	zinc
В	aqueous sodium hydroxide	zinc carbonate
С	dilute sulfuric acid	zinc
D	dilute sulfuric acid	zinc carbonate

33 Which box corresponds to limestone?



34 Petroleum is an important fossil fuel.

Which row correctly describes petroleum?

	type of substance	composition
Α	compound	mainly hydrocarbons
В	compound	only hydrogen and carbon
С	mixture	mainly hydrocarbons
D	mixture	only hydrogen and carbon

35 Butane reacts as shown.

What is this type of reaction?

- A combustion
- **B** cracking
- **C** polymerisation
- **D** reduction

36 Which substance is in the same homologous series as methanol?

- 37 Which statement could **not** be correct for an alkane?
 - **A** It burns readily in a plentiful supply of air to form only carbon dioxide and water.
 - **B** It decolourises aqueous bromine.
 - **C** It has a boiling point of -42 °C.
 - **D** The carbon and hydrogen atoms in the molecule are joined by sharing pairs of electrons.

38 In which conical flask will ethanol be produced?



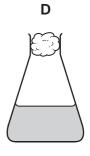
water and sugar



sugar and yeast

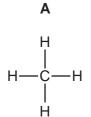


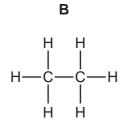
water, sugar and yeast

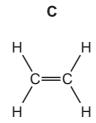


water and yeast

39 Which molecule can be polymerised?







D

40 Which row describes what happens when ethanol burns in air?

	a white powder is left	heat energy is given out	carbon dioxide is formed	water is formed
Α	✓	X	✓	✓
В	x	✓	✓	✓
С	x	✓	✓	x
D	X	✓	X	✓

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The Periodic Table of Elements

	III/	2 :	В	helium 4	10	Ne	neon 20	18	Ar	argon 40	36	궃	krypton 84	54	Xe	xenon 131	98	R	radon			
	IIA				6	ш	fluorine 19	17	Cl	chlorine 35.5	35	ğ	bromine 80	53	н	iodine 127	85	Αt	astatine -			
	I				8	0	oxygen 16	16	ഗ	sulfur 32	34	Se	selenium 79	52	<u>a</u>	tellurium 128	84	Ъ	moloud –	116	^	livermorium -
	>				7	Z	nitrogen 14	15	₾	phosphorus 31	33	As	arsenic 75	51	Sp	antimony 122	83	<u>.</u>	bismuth 209			
	<u>></u>				9	ပ	carbon 12	14	S	silicon 28	32	Ge	germanium 73	50	Sn	tin 119	82	Pb	lead 207	114	Εl	flerovium
	=				2	В	boron 11	13	Αl	aluminium 27	31	Ga	gallium 70	49	In	indium 115	81	11	thallium 204			
											30	Zu	zinc 65	48	В	cadmium 112	80	Нg	mercury 201	112	C	copernicium
											29	n	copper 64	47	Ag	silver 108	79	Au	gold 197	111	Rg	roentgenium
Group											28	Z	nickel 59	46	Pd	palladium 106	78	₫	platinum 195	110	Ds	darmstadtium -
Gro											27	ဝိ	cobalt 59	45	Rh	rhodium 103	77	Ir	iridium 192	109	Μţ	meitnerium -
		- :	I	hydrogen 1											Ru	ruthenium 101	9/	SO	osmium 190	108	Hs	hassium -
											25	M	manganese 55	43	ည	technetium -	75	Re	rhenium 186	107	Bh	bohrium —
					_	pol	ass						chromium 52		Mo	molybdenum 96	74	≥	tungsten 184	106	Sg	seaborgium -
				Key	atomic number	atomic symbo	name relative atomic mass				23	>	vanadium 51	41	g	niobium 93	73	ā	tantalum 181	105	g C	dubnium —
						ato	rek				22	F	titanium 48	40	Zr	zirconium 91	72	茔	hafnium 178	104	꿒	rutherfordium —
											21	လွ	scandium 45	39	>	yttrium 89	57–71	lanthanoids		89-103	actinoids	
	=				4	Be	beryllium 9	12	Mg	magnesium 24	20	Ca	calcium 40	38	ഗ്	strontium 88	56	Ba	barium 137	88	Ra	radium
	_				3	:=	lithium 7	=	Na	sodium 23	19	メ	potassium 39	37	Rb	rubidium 85	55	S	caesium 133	87	占	francium —

j	71	Lu	lutetium	175	103	۲	lawrencium	I	
	20	χ	ytterbium	173	102	8	nobelium	ı	
	69	Tm	thulium	169	101	Md	mendelevium	I	
;	89	ш	erbium	167	100	Fm	fermium	I	
	29	운	holmium	165	66	Es	einsteinium	ı	
	99	Ò	dysprosium	163	86	ర	californium	I	
;	65	Д	terbium	159	26	益	berkelium	I	
;	64	Вd	gadolinium	157	96	Cm	curium	I	
	63	En	europium	152	96	Am	americium	I	
	62	Sm	samarium	150	94	Pu	plutonium	ı	
	61	Pm	promethium	ı	86	ď	neptunium	ı	
1	09	PΝ	neodymium	144	92	\supset	uranium	238	
	69	Ą	praseodymium	141	91	Ра	protactinium	231	
1	28	Ce	cerium	140	06	드	thorium	232	
	22	Гa	lanthanum	139	68	Ac	actinium	ı	
		lanthanoids				actinoids			

The volume of one mole of any gas is 24 dm3 at room temperature and pressure (r.t.p.)