

CANDIDATE
NAME

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CANDIDATE
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CHEMISTRY

0620/32

Paper 3 (Extended)

May/June 2014

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 12.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

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The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **12** printed pages.

1 The table below gives the electron distributions of atoms of different elements.

element	electron distribution
A	2 + 7
B	2 + 8 + 4
C	2 + 8 + 8 + 1
D	2 + 8 + 18 + 5
E	2 + 8 + 18 + 7
F	2 + 8 + 18 + 18 + 8

For each of the following, select an element or elements from the table that matches the description. Each element may be selected once, more than once or not at all.

(a) These **two** elements are in the same group.

..... [1]

(b) This element forms a fluoride with a formula of the type XF_3 .

..... [1]

(c) This element reacts violently with cold water.

..... [1]

(d) This element has a macromolecular structure similar to that of diamond.

..... [1]

(e) The only oxidation state of this element is 0.

..... [1]

(f) This element is bromine.

..... [1]

(g) This element is a good conductor of electricity.

..... [1]

[Total: 7]

2 (a) Natural gas, which is mainly methane, is a fossil fuel.

(i) What is meant by the term *fuel*?

.....
..... [1]

(ii) Name **two** other fossil fuels.

..... [2]

(iii) Name a **solid** fuel which is not a fossil fuel.

..... [1]

(b) Fossil fuels are formed by the anaerobic decomposition of organic matter. Anaerobic means in the absence of oxygen.

(i) The organic matter contains hydrogen and carbon. Suggest the products that would be formed if the decomposition occurred in the presence of oxygen.

..... [2]

(ii) What are the **two** main disadvantages in the widespread use of fossil fuels?

.....
.....
..... [2]

[Total: 8]

3 Plant growth is improved by the availability of essential elements, such as nitrogen, and by the soil having a suitable pH.

(a) Nitrogen-based fertilisers are made from ammonia. Ammonia is manufactured by the Haber process.

(i) Describe the Haber process giving reaction conditions and a balanced equation. (Do not discuss reaction rate and yield.)

.....
.....
.....
.....
..... [5]

(ii) Fertilisers contain nitrogen. Name the other **two** elements essential for plant growth commonly found in fertilisers.

..... [2]

(b) Crops do not grow well if the soil is too acidic.

(i) One cause of acidity in soil is acid rain. Explain how acid rain is formed.

.....
.....
.....
..... [3]

(ii) Name **two** bases which are used to increase the pH of acidic soils.

..... [2]

[Total: 12]

4 Propanoic acid is a carboxylic acid. Its formula is $\text{CH}_3\text{-CH}_2\text{-COOH}$.

(a) Propanoic acid is the third member of the homologous series of carboxylic acids.

(i) Give the name and structural formula of the fourth member of this series.

name

formula [2]

(ii) Members of a homologous series have very similar chemical properties.
State **three** other characteristics of a homologous series.

.....

.....

.....

..... [3]

(b) Carboxylic acids can be made by the oxidation of alcohols.

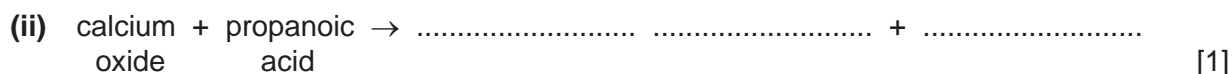
(i) Draw the structural formula of the alcohol which can be oxidised to propanoic acid.
Show all atoms and bonds.

[1]

(ii) Name a reagent, other than oxygen, which can oxidise alcohols to carboxylic acids.

..... [2]

- (c) Complete the following equations for some of the reactions of propanoic acid. The salts of this acid are called propanoates.



- (d) A piece of magnesium was added to 100 cm^3 of an aqueous acid. The time taken for the metal to react completely was measured. This experiment was repeated using different aqueous acids. The same volume of acid was used in each experiment and the pieces of magnesium used were identical. In one experiment the reaction was carried out at a different temperature.

experiment	acid	concentration in mol/dm^3	temperature / $^\circ\text{C}$	time /minutes
A	propanoic	1.0	20	5
B	propanoic	1.0	30	3
C	propanoic	0.5	20	8
D	hydrochloric	1.0	20	1

Explain the following in terms of collision rate between reacting particles.

- (i) Why is the rate in experiment **C** slower than the rate in experiment **A**?

.....

 [2]

- (ii) Why is the rate in experiment **B** faster than the rate in experiment **A**?

.....

 [2]

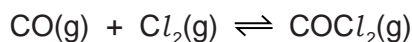
- (iii) Why is the rate in experiment **D** faster than the rate in experiment **A**?

.....

 [3]

[Total: 18]

5 Carbonyl chloride is made from carbon monoxide and chlorine.

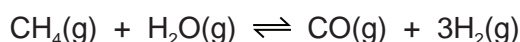


(a) Two methods of preparing carbon monoxide are from methane and oxygen, and from methane and steam.

(i) The reaction between methane and oxygen can also form carbon dioxide. How can carbon monoxide be made instead of carbon dioxide?

..... [1]

(ii) The following reaction is used to make carbon monoxide and hydrogen. The reaction is carried out at 1100 °C and normal pressure.



The reaction is reversible and comes to equilibrium. Suggest why a high temperature is used.

.....

 [2]

(iii) What is the disadvantage of using a high pressure for the reaction given in (a)(ii)?

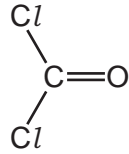
.....
 [2]

(b) Chlorine is made by the electrolysis of concentrated aqueous sodium chloride. Describe this electrolysis. Write ionic equations for the reactions at the electrodes and name the sodium compound formed.

.....

 [5]

(c) The structural formula of carbonyl chloride is given below.



Draw a diagram showing the arrangement of the valency electrons around the atoms in one molecule of this covalent compound.

Use o to represent an electron from an oxygen atom.

Use x to represent an electron from a chlorine atom.

Use • to represent an electron from a carbon atom.

[3]

[Total: 13]

6 Scandium, proton number 21, is not a typical transition element.

(a) Scandium is a low density metal which has only one oxidation state in its compounds. Scandium compounds are white solids which form colourless solutions. Titanium, the next metal in the period, is a far more typical transition element. How would the properties of titanium differ from those of scandium?

.....

.....

.....

..... [3]

(b) Scandium fluoride is an ionic compound. The valency of scandium in scandium fluoride is three.

Draw a diagram which shows the formula of this compound, the charges on the ions and the arrangement of the valency electrons around the negative ions.

Use x to represent an electron from a fluorine atom.

Use o to represent an electron from a scandium atom.

[3]

(c) Scandium oxide is insoluble in water. Describe how you could show that it is an amphoteric oxide.

.....

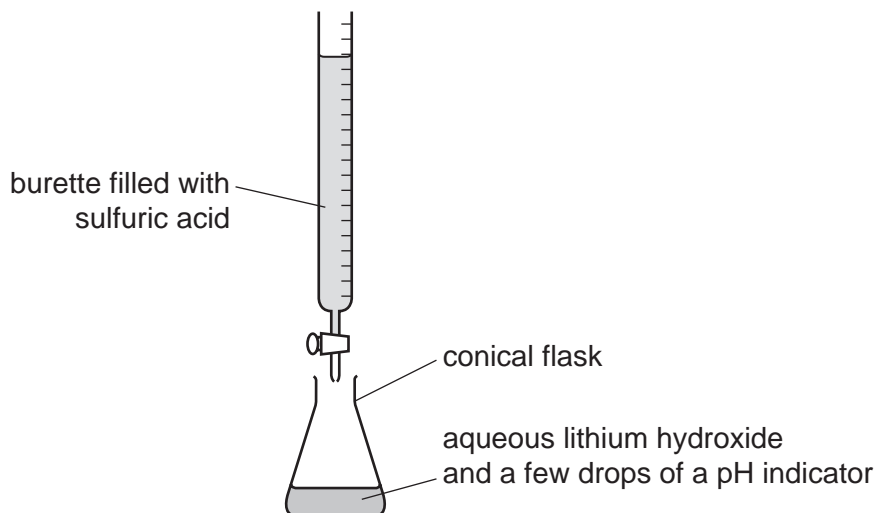
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.....

..... [3]

[Total: 9]

- 7 The soluble salt hydrated lithium sulfate is made by titration from the soluble base lithium hydroxide.



- (a) The sulfuric acid is added slowly from the burette until the indicator just changes colour. The volume of sulfuric acid needed to just neutralise the lithium hydroxide is noted. Describe how you would continue the experiment to obtain pure dry crystals of hydrated lithium sulfate.

.....

.....

.....

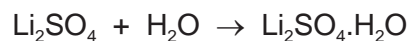
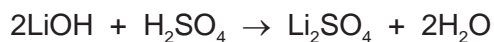
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.....

.....

..... [5]

- (b) Using 25.0 cm³ of aqueous lithium hydroxide, concentration 2.48 mol/dm³, 2.20 g of hydrated lithium sulfate was obtained. Calculate the percentage yield, giving your answer to **one** decimal place.



Number of moles of LiOH used =

Number of moles of Li₂SO₄·H₂O which could be formed =

Mass of one mole of Li₂SO₄·H₂O = 128 g

Maximum yield of Li₂SO₄·H₂O = g

Percentage yield =%

[4]

(c) An experiment was carried out to show that the formula of the hydrated salt is $\text{Li}_2\text{SO}_4 \cdot \text{H}_2\text{O}$. A sample of the hydrated salt was weighed and its mass recorded. It was then heated and the anhydrous salt was weighed. This procedure was repeated until two consecutive masses were the same. This procedure is called 'heating to constant mass'.

(i) What is the reason for heating to constant mass?

..... [1]

(ii) The mass of the hydrated salt is m_1 and the mass of the anhydrous salt is m_2 . Explain how you could show that the hydrated salt has **one** mole of water of crystallisation per mole of the anhydrous salt.

.....

.....

..... [3]

[Total: 13]

DATA SHEET
The Periodic Table of the Elements

Group		I	II	III	IV	V	VI	VII	0												
		1 H Hydrogen 1							2 He Helium 2												
3	4	7 Li Lithium	9 Be Beryllium		11 B Boron	12 C Carbon	13 Al Aluminium	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon									
11	12	23 Na Sodium	24 Mg Magnesium		27 Fe Iron	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton							
19	20	39 K Potassium	40 Ca Calcium		44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon						
37	38	85 Rb Rubidium	88 Sr Strontium		101 Ru Ruthenium	102 Rh Rhodium	103 Pd Palladium	104 Ag Silver	105 Cd Cadmium	106 In Indium	107 Sn Tin	108 Sb Antimony	109 Te Tellurium	110 I Iodine	111 Xe Xenon						
55	56	133 Cs Caesium	137 Ba Barium		186 Re Rhenium	187 Rh Rhodium	188 Pt Platinum	189 Au Gold	190 Hg Mercury	191 Tl Thallium	192 Pb Lead	193 Bi Bismuth	194 Po Polonium	195 At Astatine	196 Rn Radon						
87	88	226 Fr Francium	226 Ra Radium		227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium	227 Ac Actinium					
		*58-71 Lanthanoid series																			
		†90-103 Actinoid series																			
		Key																			
		a = relative atomic mass																			
		X = atomic symbol																			
		b = proton (atomic) number																			

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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