



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/13

Paper 1 Multiple Choice

May/June 2012

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 7 0 9 4 4 3 3 8 3 1 3 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

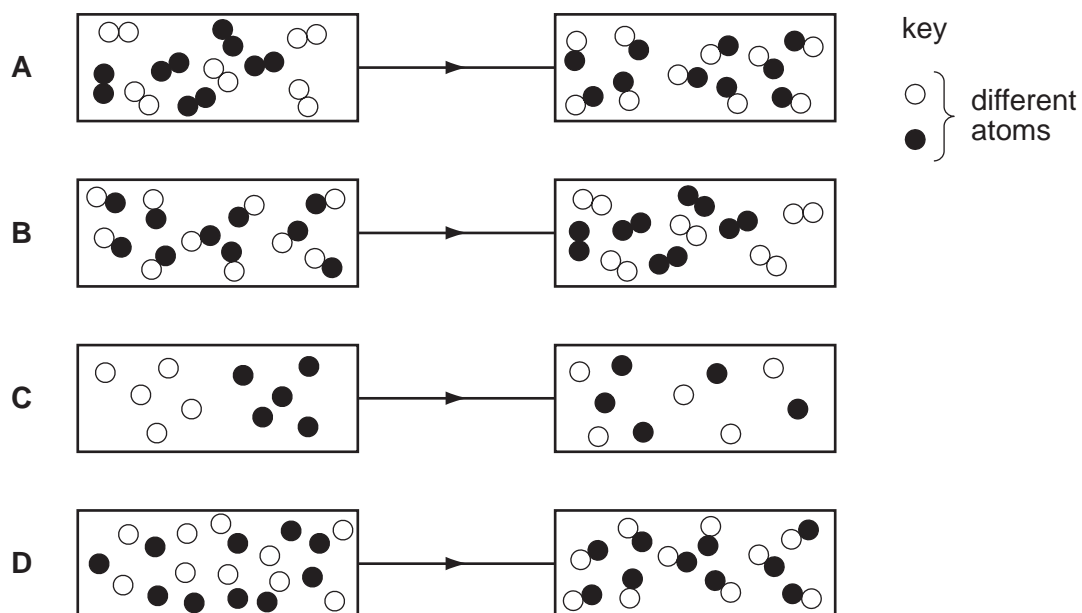
You may use a calculator.

bestexamhelp.com

This document consists of **16** printed pages.



1 Which diagram shows the process of diffusion?



2 A student investigates how the concentration of an acid affects the speed of reaction with a 0.5 g mass of magnesium at 30 °C.

The student has a beaker, concentrated acid, water and the apparatus below.

P a balance

Q a clock

R a measuring cylinder

S a thermometer

Which pieces of apparatus does the student use?

A P, Q and R only

B P, Q and S only

C Q, R and S only

D P, Q, R and S

3 Which method is most suitable to obtain zinc carbonate from a suspension of zinc carbonate in water?

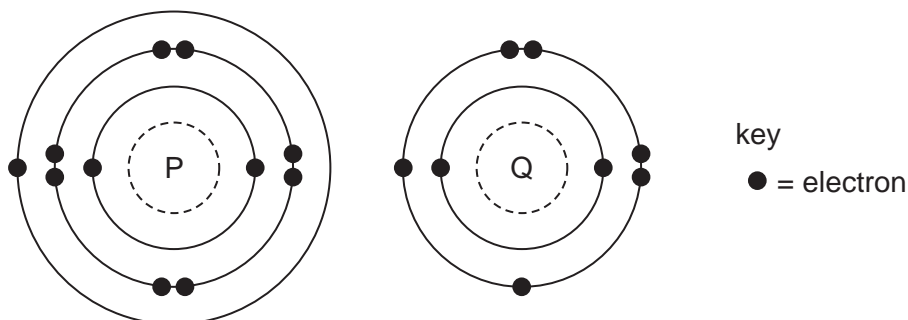
A crystallisation

B distillation

C evaporation

D filtration

- 4 The electronic structures of atoms P and Q are shown.



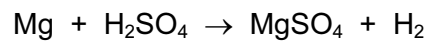
P and Q react to form an ionic compound.

What is the formula of this compound?

- A PQ_2 B P_2Q C P_2Q_6 D P_6Q_2
- 5 An element Y has the proton number 18.
- The next element in the Periodic Table is an element Z.
- Which statement is correct?
- A Element Z has one more electron in its outer shell than element Y.
- B Element Z has one more electron shell than element Y.
- C Element Z is in the same group of the Periodic Table as element Y.
- D Element Z is in the same period of the Periodic Table as element Y.
- 6 Which atom has twice as many neutrons as protons?
- A ${}^1_1\text{H}$ B ${}^2_1\text{H}$ C ${}^3_1\text{H}$ D ${}^4_2\text{He}$
- 7 Which is a simple covalent molecule?

	conducts electricity		volatile
	when solid	when molten	
A	✓	✓	x
B	✓	x	✓
C	x	✓	x
D	x	x	✓

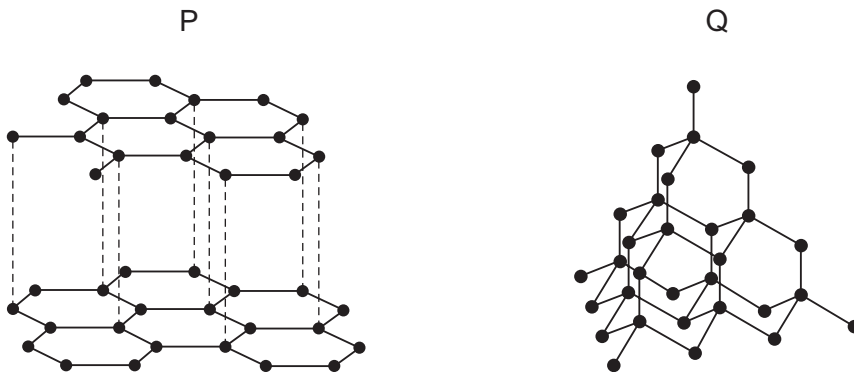
- 8 The equation for the reaction between magnesium and dilute sulfuric acid is shown.



M_r of MgSO_4 is 120

Which mass of magnesium sulfate will be formed if 12 g of magnesium are reacted with sulfuric acid?

- A 5g B 10g C 60g D 120g
- 9 The diagrams show the structures of two forms, P and Q, of a solid element.

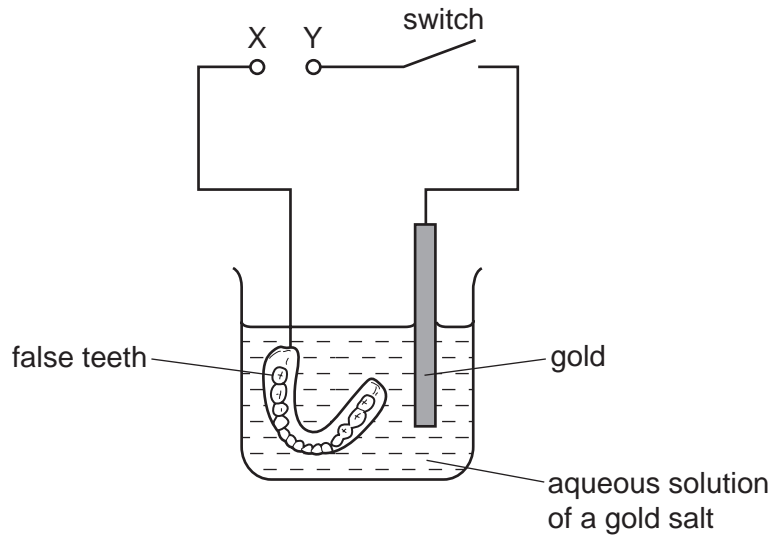


What are suitable uses of P and Q, based on their structures?

	use of solid P	use of solid Q
A	drilling	drilling
B	lubricating	drilling
C	drilling	lubricating
D	lubricating	lubricating

10 Winston Churchill, a British Prime Minister, had his false teeth electroplated with gold.

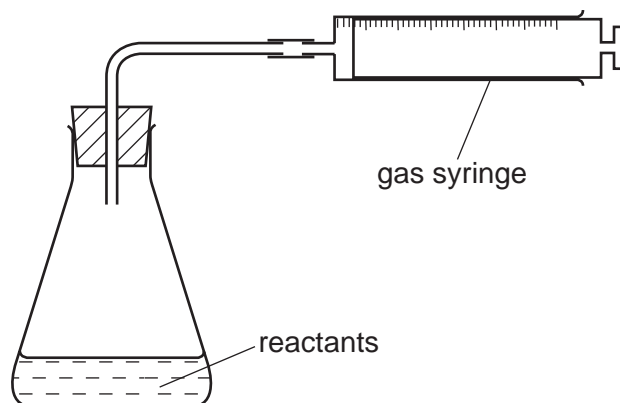
The teeth were coated with a thin layer of carbon and were then placed in the apparatus shown.



Which row is correct?

	terminal X is	the carbon powder could be
A	negative	diamond
B	negative	graphite
C	positive	diamond
D	positive	graphite

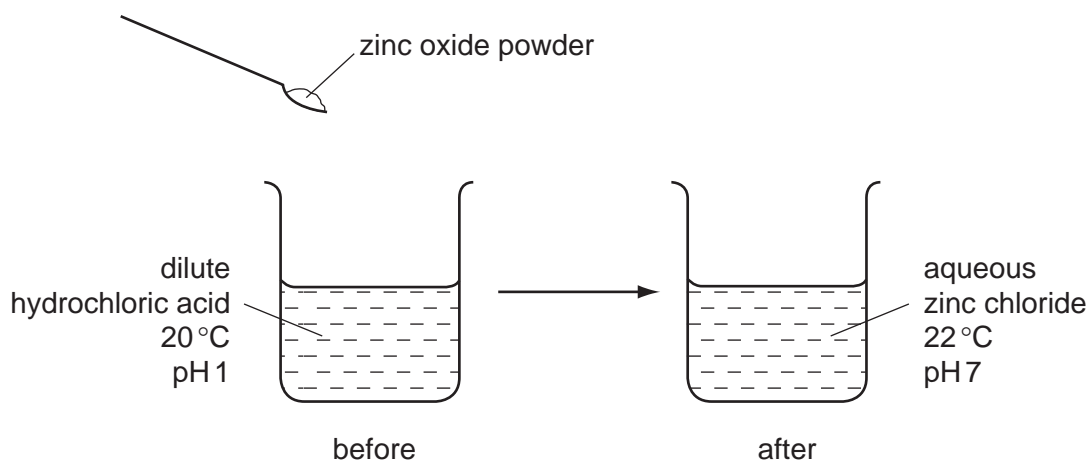
11 The apparatus shown is used to measure the speed of a reaction.



Which equation represents a reaction where the speed can be measured using this apparatus?

- A** $\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
B $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
C $\text{Fe(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{Cu(s)} + \text{FeSO}_4\text{(aq)}$
D $2\text{Na(s)} + \text{Br}_2\text{(l)} \rightarrow 2\text{NaBr(s)}$

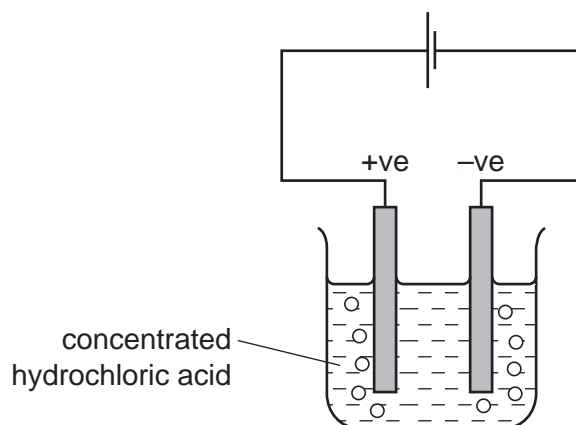
12 The diagram shows the reaction between zinc oxide and dilute hydrochloric acid.



Which terms describe the reaction?

	endothermic	neutralisation
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 13 The diagram shows that two gases are formed when concentrated hydrochloric acid is electrolysed using inert electrodes.



Which row correctly describes the colours of the gases at the electrodes?

	anode (+ve)	cathode (-ve)
A	colourless	colourless
B	colourless	yellow-green
C	yellow-green	colourless
D	yellow-green	yellow-green

- 14 A gas is escaping from a pipe in a chemical plant.

A chemist tests this gas and finds that it is alkaline.

What is this gas?

- A** ammonia
 - B** chlorine
 - C** hydrogen
 - D** sulfur dioxide
- 15 The element vanadium, V, forms several oxides.

In which change is oxidation taking place?

- A** $\text{VO}_2 \rightarrow \text{V}_2\text{O}_3$
- B** $\text{V}_2\text{O}_5 \rightarrow \text{VO}_2$
- C** $\text{V}_2\text{O}_3 \rightarrow \text{VO}$
- D** $\text{V}_2\text{O}_3 \rightarrow \text{V}_2\text{O}_5$

16 Dilute hydrochloric acid is added to a solid, S.

A flammable gas, G, is formed. Gas G is less dense than air.

What are S and G?

	solid S	gas G
A	copper	hydrogen
B	copper carbonate	carbon dioxide
C	zinc	hydrogen
D	zinc carbonate	carbon dioxide

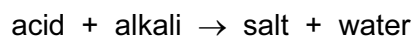
17 The results of three tests on a solution of compound X are shown in the table.

test	result
aqueous sodium hydroxide added	white precipitate formed, soluble in excess
aqueous ammonia added	white precipitate formed, insoluble in excess
acidified silver nitrate added	white precipitate formed

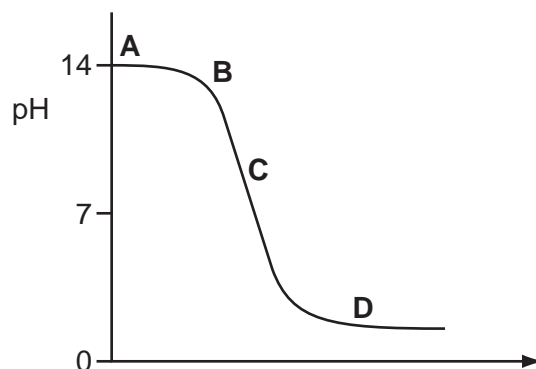
What is compound X?

- A** aluminium bromide
- B** aluminium chloride
- C** zinc bromide
- D** zinc chloride

18 The graph shows how the pH changes as an acid is added to an alkali.



Which letter represents the area of the graph where both acid and salt are present?



- 19 Which properties of the element titanium, Ti, can be predicted from its position in the Periodic Table?

	can be used as a catalyst	conducts electricity when solid	has low density	forms coloured compounds
A	✓	✓	✗	✓
B	✓	✓	✓	✗
C	✓	✗	✓	✓
D	✗	✓	✓	✓

- 20 The diagram shows a section of the Periodic Table.

Which element is described below?

'A colourless, unreactive gas that is denser than air.'

			A
	B		
		C	
			D

- 21 Element X is below iodine in the Periodic Table.

Which row correctly shows the physical state of element X at room temperature and its reactivity compared with that of iodine?

	physical state of element X at room temperature	reactivity compared with that of iodine
A	gas	less reactive
B	solid	less reactive
C	gas	more reactive
D	solid	more reactive

22 Which property is shown by **all** metals?

- A They are extracted from their ores by heating with carbon.
- B They conduct electricity.
- C They form acidic oxides.
- D They react with hydrochloric acid to form hydrogen.

23 Five elements have proton numbers 10, 12, 14, 16 and 18.

What are the proton numbers of the three elements that form oxides?

- A 10, 12 and 14
- B 10, 14 and 18
- C 12, 14 and 16
- D 14, 16 and 18

24 Metal X reacts violently with water.

Metal Y reacts slowly with steam.

Metal Z does not react with dilute hydrochloric acid.

What is the correct order of reactivity of these metals, most reactive first?

- A $X \rightarrow Y \rightarrow Z$
- B $X \rightarrow Z \rightarrow Y$
- C $Z \rightarrow X \rightarrow Y$
- D $Z \rightarrow Y \rightarrow X$

25 Which statement about the extraction of iron from its ore is correct?

- A Iron is more difficult to extract than zinc.
- B Iron is more difficult to extract than copper.
- C Iron is easy to extract because it is a transition metal.
- D Iron cannot be extracted by reduction with carbon.

26 Which statement about the uses of metals is correct?

- A Aluminium is used in the manufacture of aircraft as it has a high density.
- B Aluminium is used to make food containers as it conducts electricity.
- C Stainless steel for cutlery is made by adding other elements to iron.
- D Stainless steel is used to make chemical reactors as it corrodes readily.

27 Fertilisers need to supply crops with three main elements.

Which compound contains all three of these elements?

- A H_3PO_4 B KNO_3 C $\text{NH}_4\text{K}_2\text{PO}_4$ D NH_4NO_3

28 Some uses of water are listed.

- 1 for drinking
- 2 in chemical reactions
- 3 in swimming pools
- 4 in washing

For which uses is it necessary to chlorinate the water?

- A 1 and 2 B 1 and 3 C 2 and 4 D 3 and 4

29 Which is a use of oxygen?

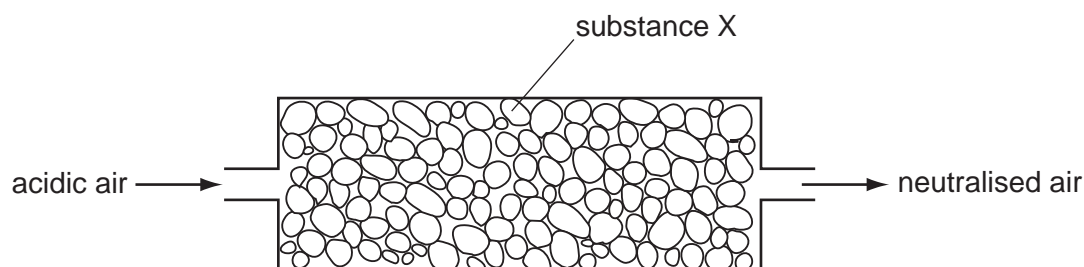
- A filling balloons
- B filling light bulbs
- C food preservation
- D making steel

30 Coal is a fossil fuel.

Which gas is **not** formed when coal burns?

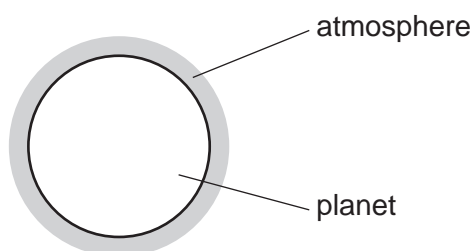
- A carbon dioxide
- B carbon monoxide
- C methane
- D sulfur dioxide

- 31 Air containing an acidic impurity was neutralised by passing it through a column containing substance X.



What is substance X?

- A calcium oxide
 - B sand
 - C sodium chloride
 - D concentrated sulfuric acid
- 32 A new planet has been discovered and its atmosphere has been analysed.



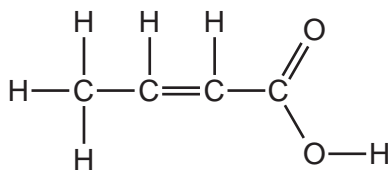
The table shows the composition of the atmosphere.

gas	percentage by volume
carbon dioxide	4
nitrogen	72
oxygen	24

Which gases are present in the atmosphere of the planet in a higher percentage than they are in the Earth's atmosphere?

- A carbon dioxide and oxygen
- B carbon dioxide only
- C nitrogen and oxygen
- D nitrogen only

33 The structure of a compound is shown.



Which functional groups are present in this compound?

	alcohol	alkene	carboxylic acid
A	✓	✓	✓
B	✓	x	x
C	x	✓	✓
D	x	x	✓

34 Gas X is a waste gas from digestion in animals.

Gas Y is formed when gas X is burnt with a small amount of oxygen.

Gas Z is formed when gas X is burnt with an excess of oxygen.

What are X, Y and Z?

	X	Y	Z
A	carbon dioxide	methane	carbon monoxide
B	carbon monoxide	methane	carbon dioxide
C	methane	carbon dioxide	carbon monoxide
D	methane	carbon monoxide	carbon dioxide

35 Which fraction from the fractional distillation of petroleum does **not** match its correct use?

	fraction	use
A	fuel oil	domestic heating
B	kerosene	jet fuel
C	naphtha	making roads
D	refinery gas	for heating and cooking

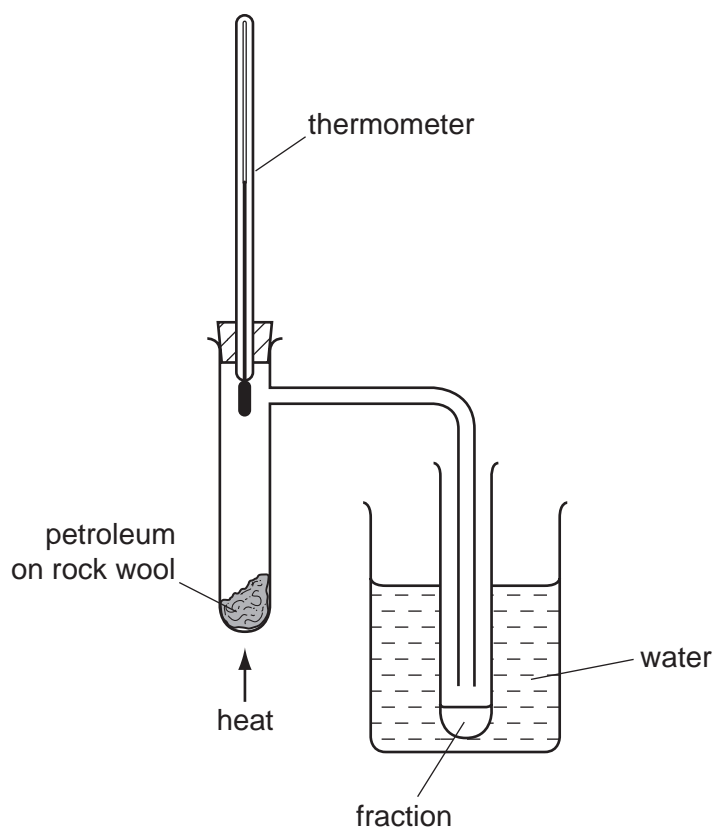
36 When a long chain hydrocarbon is cracked, the following products are produced.

- 1 C_3H_8
- 2 C_2H_4
- 3 C_3H_6
- 4 C_2H_6

Which products would decolourise bromine water?

- A** 1 and 4 **B** 2 and 3 **C** 2 only **D** 3 only

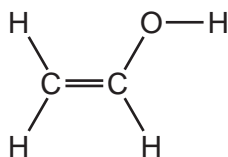
37 The diagram shows apparatus used to separate petroleum into four fractions.



Which fraction contains the smallest hydrocarbon molecules?

fraction	boiling point range / °C
A	up to 70
B	70 to 120
C	120 to 170
D	over 170

38 PVA is a polymer. The monomer has the structure shown.



To which homologous series does this compound belong?

	alcohols	alkenes
A	✓	✓
B	✓	x
C	x	✓
D	x	x

39 Ethanol is an important chemical produced by the1..... of2..... .

Which words correctly complete gaps 1 and 2?

	1	2
A	combustion	ethane
B	combustion	glucose
C	fermentation	ethane
D	fermentation	glucose

40 Which equation represents incomplete combustion of ethane?

- A** $C_2H_6 + O_2 \rightarrow 2CO + 3H_2$
- B** $C_2H_6 + 2O_2 \rightarrow 2CO_2 + 3H_2$
- C** $2C_2H_6 + 5O_2 \rightarrow 4CO + 6H_2O$
- D** $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$

DATA SHEET
The Periodic Table of the Elements

		Group										
		I	II	III	IV	V	VI	VII	VIII	IX	X	
		1 H Hydrogen 1										
		4 He Helium 2										
7	9	3	4	5	6	7	8	9	10	11	12	13
Li Lithium	Be Beryllium	B Boron	C Carbon	N Nitrogen	O Oxygen	F Fluorine	Ne Neon	Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon	P Phosphorus
11	12	13	14	15	16	17	18	19	20	21	22	23
Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon	P Phosphorus	S Sulfur	Cl Chlorine	Ar Argon	K Potassium	Ca Calcium	Sc Scandium	Ti Titanium	V Vanadium
19	20	21	22	23	24	25	26	27	28	29	30	31
K Potassium	Ca Calcium	Sc Scandium	Ti Titanium	V Vanadium	Cr Chromium	Mn Manganese	Fe Iron	Co Cobalt	Ni Nickel	Cu Copper	Zn Zinc	Ga Gallium
37	38	39	40	41	42	43	44	45	46	47	48	49
Rb Rubidium	Sr Strontium	Y Yttrium	Zr Zirconium	Nb Niobium	Mo Molybdenum	Tc Technetium	Ru Ruthenium	Rh Rhodium	Pd Palladium	Ag Silver	Cd Cadmium	In Indium
55	56	57	72	73	74	75	76	77	78	79	80	81
Cs Caesium	Ba Barium	La Lanthanum	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury	Tl Thallium
87	88	89	†	†	†	†	†	†	†	†	†	†
Fr Francium	Ra Radium	Ac Actinium										
133	137	139	178	181	184	186	190	192	195	197	201	204
Cs Caesium	Ba Barium	La Lanthanum	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury	Pb Lead
226	227	227	†	†	†	†	†	†	†	†	†	†
Fr Francium	Ra Radium	Ac Actinium										
103	102	100	99	98	97	96	95	94	93	92	91	90
Lr Lawrencium	No Nobelium	Fm Fermium	Md Mendelevium	Cf Californium	Bk Berkelium	Am Americium	Cm Curium	Pu Plutonium	Np Neptunium	U Uranium	Pa Protactinium	Th Thorium
169	173	167	169	162	159	157	152	150	144	141	140	140
Tm Thulium	Yb Ytterbium	Er Erbium	Tm Thulium	Dy Dysprosium	Tb Terbium	Gd Gadolinium	Eu Europium	Sm Samarium	Nd Neodymium	Pr Praseodymium	Ce Cerium	Ce Cerium
175	171	171	175	162	159	157	152	150	144	141	140	140
Lu Lutetium	Yb Ytterbium	Er Erbium	Tm Thulium	Dy Dysprosium	Tb Terbium	Gd Gadolinium	Eu Europium	Sm Samarium	Nd Neodymium	Pr Praseodymium	Ce Cerium	Ce Cerium
103	102	100	101	98	97	96	95	94	93	92	91	90
Lr Lawrencium	No Nobelium	Fm Fermium	Md Mendelevium	Cf Californium	Bk Berkelium	Cm Curium	Am Americium	Pu Plutonium	Np Neptunium	U Uranium	Pa Protactinium	Th Thorium

*58-71 Lanthanoid series
†90-103 Actinoid series

a = relative atomic mass

X = atomic symbol

b = proton (atomic) number

Key

a	X	b		

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.