
CHEMISTRY

9701/35

Paper 35 (Advanced Practical Skills 1)

May/June 2018

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

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Question	Answer	Marks
1(a)	I Three masses recorded and two burette readings and titre for the rough titration and initial and final readings for two (or more) accurate titrations (<i>minimum 2 × 2 box</i>)	1
	II All three headings and units correct for accurate titrations Headings: initial / final (burette) and reading / volume / vol or reading / volume / vol at start / finish (but not V) and volume / FA 2 and added / used or titre and Units: (cm ³) or / cm ³ or in cm ³ [or cm ³ by every entry]	1
	III All accurate burette readings are recorded to the nearest 0.05 cm ³ <i>The requirement to record to 0.05 applies to burette readings, including 0.00 (if this was the initial reading) but it does not apply to the titre.</i> <i>Do not award this mark if:</i> <ul style="list-style-type: none"> • 50(.00) is used as an initial burette reading; • more than one final burette reading is 50(.00); • any burette reading is greater than 50(.00); • there is only one accurate titration 	1
	IV The final accurate titre recorded is within 0.10 cm ³ of any other accurate titre. <ul style="list-style-type: none"> • <i>Do not award the mark if any 'accurate' burette readings (apart from initial 0) are given to zero dp.</i> 	1
<p>For assessment of accuracy (Q marks) the examiner should round any burette reading to the nearest 0.05 cm³. Subtractions should be checked. The 'best' titres should be selected using the hierarchy: two (or more) identical; then 2 (or more) within 0.05 cm³; then two (or more) within 0.1 cm³ etc. These best titres should be used to calculate the mean corrected titre to the nearest 0.01 cm³. Calculate mean titre/mass used to 2 dp, calculate the supervisor titre/mass, calculate the difference, δ, and award accuracy (Q) marks as shown below.</p>		
1(a)	V, VI and VII Award V, VI and VII for a difference from supervisor within 0.10 Award V and VI for a difference from supervisor $0.10 < \delta \leq 0.20$ Award V for a difference from supervisor $0.20 < \delta \leq 0.30$	3

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Question	Answer	Marks
1(b)	<p>Candidate must average two (or more) for which the total spread is not greater than 0.2 cm³. Working must be shown or ticks must be put next to the two (or more) accurate readings selected. <i>The mean should normally be quoted to 2 dp rounded to the nearest 0.01. Example: 26.667 must be rounded to 26.67.</i> <i>Two special cases where the mean may not be to 2 dp:</i> <i>allow mean to 3 dp only for 0.025 or 0.075 e.g. 26.325;</i> <i>allow mean to 1 dp if all accurate burette readings were given to 1 dp and the mean is exactly correct. e.g. 26.0 and 26.2 = 26.1 is correct but 26.0 and 26.1 = 26.1 is incorrect.</i> <i>Do not award this mark if:</i> <i>any selected titre is not within 0.20 cm³ of any other selected titre;</i> <i>the rough titre was used to calculate the mean;</i> <i>the candidate carried out only 1 accurate titration;</i> <i>burette readings were incorrectly subtracted to obtain any of the accurate titre values.</i> <i>All burette readings, excluding initial 0, (resulting in titre values used in calculation of mean) are integers.</i></p>	1
	<p><i>Note: the candidate's mean will sometimes be marked as correct even if it is different from the mean calculated by the examiner for the purpose of assessing accuracy.</i></p>	
1(c)(i)	Answers to (c)(ii)–(iv) correct to 3 or 4 sig figs. (minimum of 3 answers attempted)	1
1(c)(ii)	Correctly calculates moles acid = (b) × 0.110 / 1000	1
1(c)(iii)	<p>Na₂CO₃(aq) + 2HCl(aq) → 2NaCl(aq) + CO₂(g) + H₂O(l) and correctly uses moles sodium carbonate = (c)(ii) ÷ 2</p>	1
1(c)(iv)	<p>Correctly uses (c)(iii) × 10 and M_r = mass from (a) ÷ (answer to (c)(iii) × 10)</p>	1
1(c)(v)	<p>$\frac{M_r - 106}{18}$ and answer as integer from working</p>	1
1(d)	Percentage error increases as volume is smaller.	1
	Half the volume of sulfuric acid is needed	1

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Question	Answer	Marks
Question	Answer	Marks
2(a)	Data recorded (in table) showing all headings and units <ul style="list-style-type: none"> • Mass of crucible + lid • Mass of crucible + lid + FA 4 • Mass of crucible + lid + residue / anhydrous FA 4 • Mass of FA 4 • Mass residue / anhydrous FA 4 • Mass water (lost) and the 3 balance readings clearly recorded to consistent number of dp. <i>Do not allow weight.</i>	1
	Mass between 1.80 g and 2.00 g used and masses water, hydrated and anhydrous solid correctly calculated	1

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Question	Answer	Marks
For assessment of accuracy (Q) marks the examiner should check and correct the masses of water and anhydrous solid and calculate ratio of mass water/mass anhydrous solid correct to 2 dp. (theoretical value 1.05)		
2(a)	Award 2 marks if ratio 0.95–1.10 Award 1 mark if ratio 0.80–1.15	2
2(b)(i)	Correctly calculates: mass water for (a) ÷ 18, to 2 – 4 sf	1
2(b)(ii)	Correctly uses (b)(i) ÷ 7 and answer to 2 – 4 sf	1
2(b)(iii)	Expression: $M_r = \text{mass anhydrous solid} \div \mathbf{(b)(ii)}$	1
	Expression: $A_r = M_r - 96.1$	1
2(b)(iv)	Identification of Y as Group 2 element with A_r nearest to (b)(iii)	1
2(c)	Moles / mass water too small / not all water lost so A_r too large	1
 as moles YSO_4 too small	1

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Question	Answer	Marks
FA 4 is MgSO₄; FA 5 is FeSO₄.7H₂O		
3(a)(i)	Chooses NaOH and NH ₃	1
	Uses both reagents in excess	1
	White ppt and insoluble in excess for both	1
3(a)(ii)	If Mg ⁽²⁺⁾ in Q2 then (observations) confirm it (owtte) or If other cation / metal identified in Q2 then not confirmed as (FA 4 contains) Mg ²⁺ / magnesium (owtte)	1
3(b)(i)	Observations (4 possible): <ul style="list-style-type: none"> • Water vapour / condensation • Gas turns litmus red • White smoke (not white fumes / white gas) • Brown / yellow / grey residue Any 3 obs = 2 marks, any 2 obs = 1 mark	2

Question	Answer		Marks
3(b)(ii)	<i>test</i>	<i>observations</i>	6
	+ NaOH	Green ppt and insol in excess / turns brown [1]	
	+ KMnO ₄	Purple to colourless / yellow [1]	
	+ NH ₄ SCN	Red / blood-red / blood-orange (solution) [1]	
	+H ₂ O ₂	Fizz and gas relights glowing splint [1]	
	then + NaOH	Brown / rust / red-brown ppt and insol in excess [1]	
	+ AgNO ₃	No visible reaction / no change / no ppt (ignore faint white ppt) and	
	+ Ba(NO ₃) ₂	White ppt and	
	then + HCl	(ppt) insoluble [1]	
3(b)(iii)	Fe ²⁺ and SO ₄ ²⁻		1
3(b)(iv)	redox / Fe ²⁺ is oxidised / MnO ₄ ⁻ is reduced		1