

CANDIDATE
NAME

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NUMBER

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CHEMISTRY

0620/32

Paper 3 Theory (Core)

October/November 2016

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

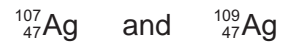
A copy of the Periodic Table is printed on page 20.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

(b) Silver has two naturally occurring isotopes.



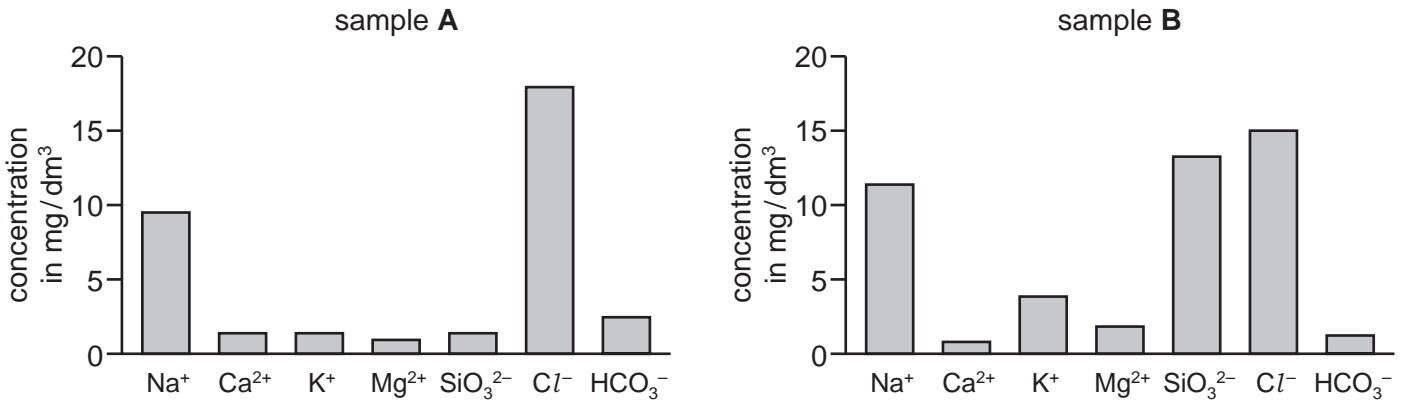
Complete the table to show the number of protons, electrons and neutrons in these **two** isotopes.

	${}_{47}^{107}\text{Ag}$	${}_{47}^{109}\text{Ag}$
number of protons		
number of electrons		
number of neutrons		

[3]

[Total: 8]

2 The bar charts compare the concentrations of ions in two samples of water, sample **A** and sample **B**.



(a) Use the information in the bar charts to answer the following questions.

(i) Describe **two** differences in the composition of sample **A** and sample **B**.

.....

 [2]

(ii) Which positive ion has the lowest concentration in sample **B**?

..... [1]

(iii) Calculate the mass of chloride ions present in 100 cm³ of sample **B**.
 Show all your working. [1 dm³ = 1000 cm³]

mass = mg [2]

(b) Describe a test for chloride ions.

test

result

[2]

- (c) River water contains small particles of clay.
These particles show Brownian motion.

Which **one** of these statements best describes Brownian motion?

Tick **one** box.

the diffusion of gases

the random movement of particles in a suspension

the downward movement of particles in a suspension

[1]

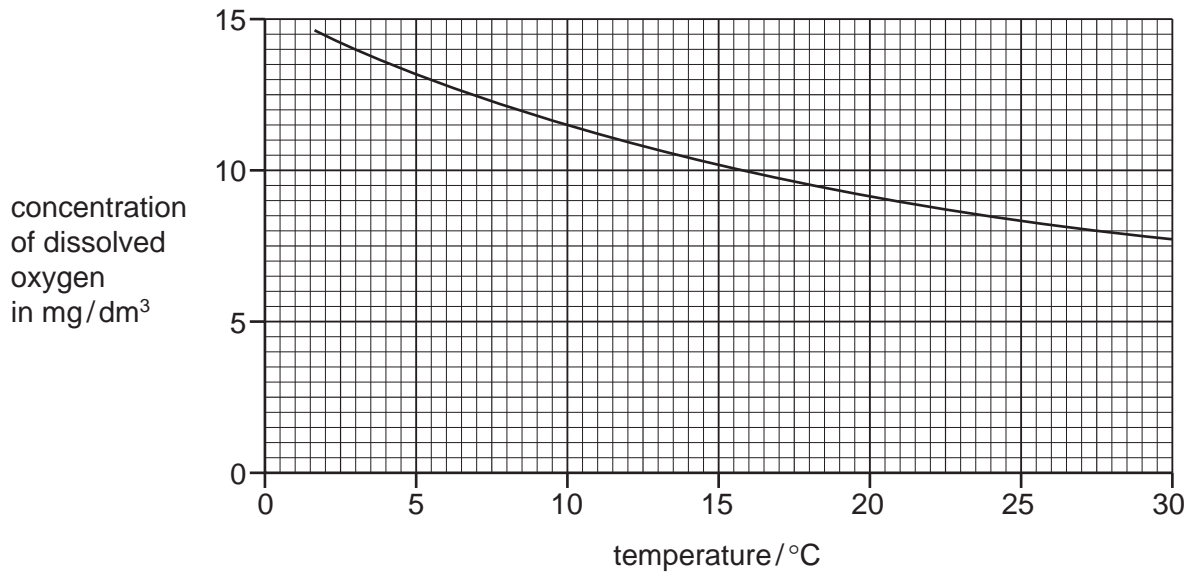
- (d) Silicon in river water comes from silicate rocks. Some of these contain silicon(IV) oxide.

Explain why silicon(IV) oxide is an acidic oxide.

..... [1]

(e) River water contains dissolved oxygen.

The graph shows how the concentration of dissolved oxygen changes with temperature.



(i) Describe how the concentration of dissolved oxygen changes with temperature.

..... [1]

(ii) Determine the concentration of oxygen present in the water at 10 °C.

..... [1]

(iii) Suggest how the rate of corrosion of iron water pipes changes with temperature.
Explain your answer.

.....
..... [1]

(f) Describe how water is treated to make it suitable to drink.

.....
..... [2]

(g) Oxides of nitrogen are common pollutants in the air.

(i) State **one** source of oxides of nitrogen in the air.

..... [1]

(ii) State **one** adverse effect of oxides of nitrogen on health.

..... [1]

[Total: 16]

3 Iron is a metal.

(a) The equation for the reaction of iron with steam is shown.



Which substance is reduced in this reaction?

Explain your answer.

.....
 [2]

(b) Iron is extracted by heating iron ore with carbon in a blast furnace.

(i) What is the meaning of the term *ore*?

..... [1]

(ii) Air is blown into the blast furnace.

What is the purpose of this air?

..... [1]

(iii) The impurities in the iron ore are removed as slag.

Which **one** of the following is slag?

Tick **one** box.

iron(II) oxide

calcium silicate

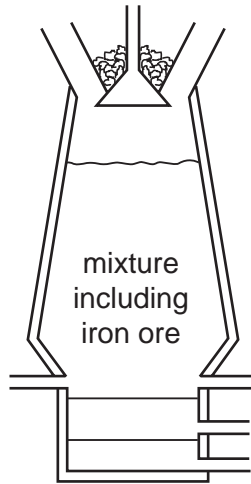
calcium carbonate

coke

[1]

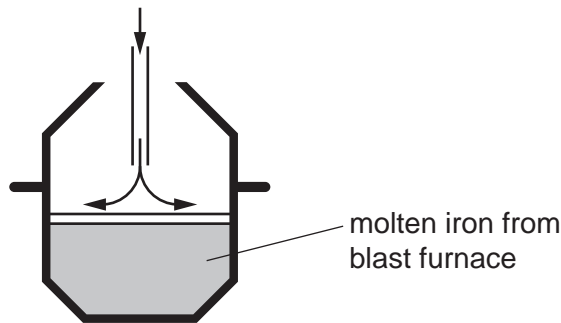
- (iv) Slag is less dense than iron.
The diagram shows a blast furnace used to extract iron.

On the diagram, write the letter **S** to show where the slag is removed.



[1]

- (c) Iron from the blast furnace contains impurities.
The diagram shows a converter used to make steel from iron.



Describe how iron is converted into steel.
In your answer

- describe the impurities present,
- describe how the impurities are removed,
- include a relevant word equation.

.....

.....

.....

.....

.....

..... [4]

[Total: 10]

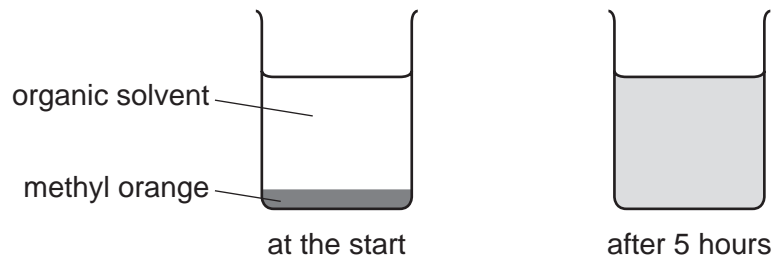
4 Methyl orange and methyl red are both dyes which can be used as indicators.

- (a) The actual value for the melting point of methyl red is 180°C .
A chemist prepares a sample of methyl red and finds that it melts over the range 173°C to 177°C .

Suggest why the melting point of this sample was different from the actual value.

..... [1]

- (b) A concentrated solution of methyl orange was placed at the bottom of a beaker containing an organic solvent.
After 5 hours, the orange colour had spread throughout the solvent.



Use the kinetic particle model of matter to explain this observation.

.....

 [3]

- (c) Methyl orange is used as an indicator.

What colour is methyl orange when placed in dilute sulfuric acid?

..... [1]

(d) Sulfuric acid can be used to prepare copper(II) sulfate from copper(II) oxide.

(i) Complete the general word equation for this reaction.

metal oxide + acid → +

[2]

(ii) Sulfuric acid is added to excess copper(II) oxide. The mixture is heated and the unreacted copper(II) oxide is removed.

Suggest how the unreacted copper(II) oxide is removed.

..... [1]

(iii) Put statements **A** to **E** about the preparation of pure dry crystals of copper(II) sulfate from copper(II) sulfate solution in the correct order.

- A** The crystals are filtered off.
- B** The heating is stopped when the point of crystallisation is reached.
- C** The mixture is left to form crystals.
- D** The crystals are dried with filter paper.
- E** The solution is heated gently.

correct order

--	--	--	--	--

[2]

[Total: 10]

- 5 Cement is made by heating clay with limestone. Some of the limestone (calcium carbonate) breaks down to form calcium oxide and a gas which turns limewater milky.

(a) (i) Complete the chemical equation for this reaction.



[2]

(ii) What type of chemical reaction is this?

..... [1]

(iii) Determine the relative formula mass of calcium carbonate. Show all your working.

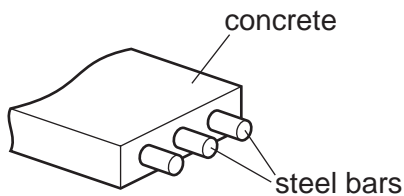
[2]

(b) Concrete is a mixture of cement, sand, water and small stones.
Calcium carbonate is a compound, but concrete is a mixture.

State **two** differences between a compound and a mixture.

.....
.....
.....
..... [2]

(c) Reinforced concrete contains steel bars within the concrete.



Some properties of concrete and steel are shown in the table.

	relative strength	relative expansion when heated	relative heat conductivity	cost
concrete	60	12	1.5	low
steel	250	12	60.0	high

Use the information in the table to suggest why concrete must be reinforced with steel when it is used to make bridges.

.....
 [1]

(d) If reinforced concrete becomes cracked, liquids and gases can reach the steel bars. The steel bars rust.

Which **two** substances are needed for steel to rust?

..... and [2]

[Total: 10]

6 Petroleum can be separated into useful hydrocarbon fractions by fractional distillation.

(a) (i) Explain the term *hydrocarbon fraction*.

hydrocarbon

fraction

..... [2]

(ii) State **one** use for each of the following hydrocarbon fractions.

naphtha

kerosene

[2]

(b) Organic compounds can be grouped into different homologous series.

Explain the term *homologous series* by referring to alkenes.

.....
.....
.....
.....
.....
..... [4]

(c) The table shows some information about alkenes.

alkene	formula	density of liquid alkene in g/cm ³	melting point /°C	boiling point /°C
ethene	C ₂ H ₄	0.568	-169	-104
propene	C ₃ H ₆	0.610	-185	-47
butene	C ₄ H ₈	0.626	-185	-6
pentene	C ₅ H ₁₀	0.640	-165	+30
hexene	C ₆ H ₁₂	0.673	-140	

- (i) A student predicts that the density of the liquid alkenes increases as the number of carbon atoms increases.

Describe whether the data in the table support this prediction.

.....
 [1]

- (ii) Predict the boiling point of hexene.

..... [1]

- (iii) Deduce the state of pentene at -60 °C.
 Explain your answer.

.....
 [2]

(d) Draw the structure of ethene. Show all of the atoms and all of the bonds.

[1]

- (e) Alkenes are manufactured by cracking.

When tetradecane, C₁₄H₃₀, is cracked the products are ethene, an alkene with four carbon atoms and an alkane.

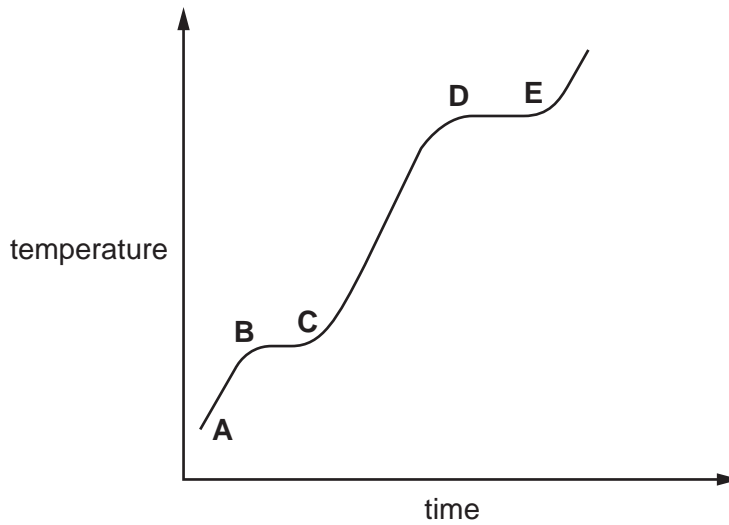
Complete the chemical equation for this reaction.



[2]

[Total: 15]

7 The graph shows how the temperature of sodium changes when it is heated at a constant rate in an atmosphere of argon.



(a) Suggest why the sodium is heated in argon and **not** in air.

..... [1]

(b) Which part of the graph above represents the boiling point of sodium?
Tick **one** box.

- A-B
- B-C
- C-D
- D-E

[1]

(c) (i) Describe **two** differences in the general properties of a liquid and a gas.

.....

 [2]

(ii) Describe the arrangement and motion of the particles in a liquid.

arrangement

motion

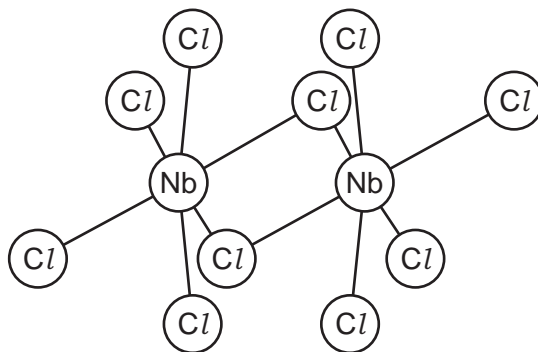
[2]

(d) Niobium is a transition element. Sodium is an element in Group I of the Periodic Table.

(i) Describe **two** properties of niobium which are different from sodium.

.....
 [2]

(ii) The structure of niobium chloride is shown.



Determine the formula of niobium chloride.

..... [1]

(iii) Niobium chloride is a covalent molecule.

Predict **two** physical properties of niobium chloride.

.....
 [2]

[Total: 11]

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The Periodic Table of Elements

		Group																																	
I	II	III	IV	V	VI	VII	VIII																												
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																		
Li lithium 7	Be beryllium 9	B boron 11	C carbon 12	Al aluminium 13	Si silicon 14	P phosphorus 15	S sulfur 16	Cl chlorine 17	Ar argon 18	K potassium 19	Ca calcium 20	Sc scandium 21	Ti titanium 22	V vanadium 23	Cr chromium 24	Mn manganese 25	Fe iron 26	Co cobalt 27	Ni nickel 28	Cu copper 29	Zn zinc 30	Ga gallium 31	Ge germanium 32	As arsenic 33	Se selenium 34	Br bromine 35	Kr krypton 36								
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57-71 lanthanoids	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 90	Nb niobium 91	Mo molybdenum 92	Tc technetium 93	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106	Ag silver 108	Cd cadmium 112	In indium 115	Sn tin 119	Sb antimony 122	Te tellurium 128	I iodine 127	Xe xenon 131	Cs caesium 133	Ba barium 137	La lanthanum 139	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195	Au gold 197	Hg mercury 201	Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium 210	At astatine 210	Rn radon 222
87	88	89-103 actinoids	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136
Fr francium	Ra radium	Ac actinium	Rf rutherfordium	Db dubnium	Sg seaborgium	Bh bohrium	Hs hassium	Mt meitnerium	Ds darmstadtium	Rg roentgenium	Cn copernicium	Nh nihonium	Fl flerovium	Lv livermorium	Ts tennessine	Og oganeson	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136

1
H
hydrogen
1

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium	94 Pu plutonium	95 Am americium	96 Cm curium	97 Bk berkelium	98 Cf californium	99 Es einsteinium	100 Fm fermium	101 Md mendelevium	102 No nobelium	103 Lr lawrencium

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).